

Instructions.

Cut out the properties on the first sheet and place them in the correct box on the second sheet.

Have very high melting points
Very good conductor of electricity
Dents when hammered
Made of molecules or atoms
Made of anions and cations
Made of metal atoms
Made of molecules or atoms
Held together by weak Van der Waals forces
Held together by strong electrostatic attraction
Held together by strong metallic bonding
Held together by strong covalent bonding
Diamond, graphite and silica are examples
Do not dissolve in any solvent
Does not conduct electricity (no charged particles)
Very hard and brittle
Consist of a regular arrangement (array) of positive ions in a sea of mobile electrons
Have fairly high melting points
Are malleable
Do not dissolve in any solvent (bonds are too strong to be broken by the attraction of any solvent) unless they react with water

Consist of a regular array(arrangement) of positive and negative ions
CaO, Ag NO ₃ , CuSO ₄ , Ca CO ₃ are examples
Zinc, aluminium, iron and copper are examples
Mercury is the only liquid example
Have low melting points and some sublime
Do not conduct electricity unless dissolved or molten
Usually dissolve in non-polar solvents
Are easily broken or cut with a knife
I ₂ , Br ₂ and CO ₂ are examples
Usually do not conduct electricity

Thanks to Jane Sanfelieu (Sacred Heart - NP)

Molecular Solids

Ionic Solids

Metallic Solids

Extended Covalent Solids